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temperature of 20

° C. The specific heat of A is larger than that of B. Under the

same heating

conditions, which

metal would take

longer to reach 21

° C? Explain your

reasoning. 2.

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step problem-solving  
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kilojoules of energy  
are in a donut that  
contains 200.0  
Calories? 2. What is  
the specific heat of a  
substance that has a  
mass of 25.0 g and  
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Heat Problems 1)  
How much heat must  
be absorbed by 375  
grams of water to  
raise its temperature  
by  $25^{\circ}\text{C}$ ? 2) What  
mass of water can be

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heated from  $25.0^{\circ}\text{C}$   
to  $50.0^{\circ}\text{C}$  by the  
addition of  $2825\text{ J}$ ?

3) What is the final  
temperature when  
 $625\text{ grams}$  of water  
at  $75.0^{\circ}\text{C}$  loses  $7.96$   
 $\times 10^4\text{ J}$ ? Chapter 17.

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A 12.2 g sample of an unknown metal sample was heated to  $98.6\text{ }^{\circ}\text{C}$  and then put into a calorimeter containing 25.0 ml of water at  $22.3\text{ }^{\circ}\text{C}$ .

After mixing, the temperature of the

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water and metal  
increased to a  
maximum of  $28.4^{\circ}\text{C}$   
after 35 seconds.

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